

Process for the preparation of methylheptenone

The present invention relates to a novel process for the preparation of 2-methyl-2-hepten-6-one, hereinafter referred to also as methylheptenone. It is known to prepare methyl-
5 heptenone by reaction of 2-methyl-3-buten-2-ol with isopropenyl methyl ether in the presence of catalysts, e.g phosphoric acid, see DE 1 193 490 or organic phosphates or phosphinates, see DE 196 49 564.

It has been found that hydrogen bis(oxalato)borate and, particularly, hydrogen
10 tris(oxalato)phosphate are especially well suited to catalyze the reaction. Accordingly, the present invention is concerned with a process for the preparation of 2-methyl-2-hepten-6-one which comprises reacting 2-methyl-3-buten-2-ol with isopropenyl methyl ether in the presence of hydrogen tris(oxalato)phosphate or hydrogen bis(oxalato)borate.

15 The process of the present invention can be carried out at normal pressure or at elevated pressure. Preferably, the reaction is carried out at elevated pressure, e.g at a pressure of about 10^5 to about 20×10^5 Pascal (Pa), preferred 5×10^5 to about 15×10^5 Pa. Further, the reaction according to the present invention is preferably carried out in the absence of a solvent. The catalyst can be added in pure form or solved, preferred is the solved form in
20 acetone or methanol. The reaction is suitably carried out at elevated temperature, e.g. at a temperature between about 373 K and about 450 K. The molar ratio of the reactants is suitably about 1 : 1 to 1 : 3, preferred is 1 : 2 to 1 : 2.5 although an excess of one of the reactants, e.g., an excess of the less expensive reactant, isopropenyl methyl ether, may be used and, if desired, be recycled after isolation of the desired reaction product from the
25 reaction mixture.

The reaction in accordance with the present invention can be carried out batchwise or in continuous mode, e.g., in two or more serially arranged stirred tanks or in tube reactors.

The catalysts, hydrogen tris(oxalato)phosphate or hydrogen bis(oxalato)borate, may be present in the reaction mixture in amount to provide a substrate/catalyst ratio of about 1000:1 to about 100:1, preferably about 900:1 to about 300:1, based on 2-methyl-3-buten-2-ol as the substrate. Hydrogen tris(oxalato)phosphate can be prepared as disclosed in EP 1 227 089. Hydrogen bis(oxalato)borate can be prepared by the procedure described in EP 784 042.

The invention is illustrated further by the Example which follows.

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Example

The reaction was carried out using an equipment the main part was a stainless steel batch reactor (Medimex – High Pressure) with a nominal volume of 1.0 liter, an operative temperature up to 493 K and a maximum pressure of 20 bar. The reactor was equipped with a heating system (electrical heating spirals located in the jacket), a cooling system (glycol-water mixture flowing through the reactor jacket), a temperature control for the measurement of the inner reactor, jacket and cooling temperatures, and a stirrer. The reactor was coupled with a laboratory vacuum pump to remove the air from the reactor. The sampling was done via a stainless steel capillary and a thin spiraled stainless steel tube connected to the sampling flask. The catalysts were used as about 12 or 5.5 wt.-% solution in methanol or 6.5 wt.-% solution in acetone (hydrogen bis(oxalato)borate), or about 12, 10 or 8 wt.-% solution in methanol or 6 wt.-% solution in acetone (hydrogen tris(oxalato)phosphate).

25 195 g MBE (2.22 mol), catalyst (see Table 1) and 342 g IPM (4.6 mol = 2.1 eq.) were mixed and charged into the reactor. The reactor was shortly evacuated with a vacuum pump. The reaction mixture was heated up to 423 K while stirring and maintaining said temperature. Samples were taken in appropriate time intervals through the sample valve and capillary and were directly neutralized with sodium acetate. The samples were analyzed by gas chromatography (NORM%). After the reaction mixture was cooled to room temperature a sample of the remaining reaction mixture was taken (no gas phase) and analyzed by gas chromatography (NORM%).

The results are shown in Tables 1, 2 and 3 below.

Table 1

Catalyst (mol-%) ¹	Catalyst solution	T (K)	Reaction time (h)	Conversion of MBE (%)	Yield MH (%)
HOB (0.15)	5.5 % methanol	423	21.7	25.0	23.9
HOB (0.10)	5.5% methanol	423	24	29.9	29.5
HOB (0.10)	6.5 % acetone	423	24	48.0	23.3
HOB (0.05)	6.5 % acetone	423	24	35.7	20.1
HOB (0.01)	6.5 % acetone	423	24	16.6	14.4
HOP (0.15)	10.0 % methanol	423	21.3	96.3	91.0
HOP (0.15)	11.7 % methanol	423	24	97.7	92.5
HOP (0.15)	11.7 % methanol	423	24	97.9	97.7 ³
HOP (0.15)	11.7 % methanol	423	24	97.6	97.6 ⁴
HOP (0.12)	10.0% methanol	423	21.5	96.7	80.6
HOP (0.09)	10.0 % methanol	423	17	81.5	80.8
HOP (0.23)	10.0 % methanol	398	25	87.1	80.2
HOP (0.17)	10.0 % methanol	398	22	77.5	77.3 ²
HOP (0.15)	6.1 % acetone	423	24	94.9	56.4

Table 2

HOB (0.9)	11.6 % methanol	423	24	50.3	43.2
HOB (0.3)	11.6 % methanol	423	24	22.0	20.3 ⁴
HOB (0.3)	11.6 % methanol	423	24	28.5	26.1 ³
HOB (0.3)	11.6 % methanol	448	24	39.2	37.6
HOB (0.3)	5.5 % methanol	423	24	29.8	28.0

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Table 3

HOP (0.14)	6.1 % acetone	423	24	98.6	98.6 ⁴
HOP (0.30)	11.7 % methanol	423	8.5	94.0	90.1
HOP (0.30)	7.9 % methanol	398	24	88.3	83.4
HOP (0.30)	11.7 % methanol	423	24	99.0	81.1

HOB : hydrogen bis(oxalato)borate; HOP : hydrogen tris(oxalato)phosphate;

MH : methylheptenone; MBE : 2-methyl-3-buten-2-ol; IPM : isopropenyl methyl ether

10 ¹ based on MBE

² ratio IPM/MBE = 3.3/1 (all other cases 2.1/1), ³ ratio IPM/MBE = 2.5/1, ⁴ ratio IPM/MBE = 3.0/1